

Roll No. _____

End Semester Examination (December 2024)

B. Pharm. – 1st Semester

Subject Code: BP104T

Subject Name: Pharmaceutical Inorganic Chemistry

Time: 3.00 Hour

Total Marks: 75

		Marks	CO	PO
Q.1	Attempt all Questions	20(10*2)		
a.	What is the use of Dil. HCl in limit test of sulphate?	2	1	6
b.	Chemical formula for potash alum is _____.	2	5	6
c.	Ferrous sulphate is used as _____	2	5	6
d.	Which of the following is used as antimicrobials? 1. Hydrogen peroxide 2. Chlorinated lime 3. Potassium permanganate 4. All of the above	2	3	2
e.	Radiopharmaceutical products are stored in 1. Refrigerator 2. Glass container 3. Lead Container 4. Both 2 and 3	2	4	2
f.	Define Radiopharmaceutics.	2	4	2
g.	Give the reaction involved in limit test of iron	2	1	6
h.	Write the principle involved in limit test of chloride	2	1	6
i.	Give the uses of Boric acid	2	3	2
j.	Define limit test.	2	1	2
Q.2	Attempt any Two Questions	20(2*10)		
a.	Define Antimicrobials and explain Hydrogen peroxide	10	3	6
b.	Write a short note on application, storage and precaution of Radiopharmaceutical products.	10	4	2
c.	Explain modified limit test of chloride	10	1	6
Q.3	Attempt any Seven Questions	35(7*5)		
a.	Give uses preparation and Storage condition of Boric acid	5	5	6
b.	What is meant by buffered isotonic solutions?	5	2	2
c.	What do mean by antacids. Explain any one compound	5	3	6
d.	Write a short note on Acid-Base theory.	5	2	2
e.	Give uses preparation and Storage condition of Chlorinated lime	5	5	4
f.	Write a short note on limit test of sulphate	5	1	6
g.	Explain Haematinics and write a note on ferrous Sulphate.	5	5	6
h.	Explain various Electrolytes used in the replacement therapy	5	2	2
i.	Write a note on different Sources of impurities	5	1	2

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End Semester Examination (December 2024)
B. Pharm. – 1st Semester

Subject Code: BP102T

Subject Name: Pharmaceutical Analysis-I

Time: 3.00 Hour

Total Marks: 75

		Marks	CO	PO
Q.1	Attempt all Questions	20(10x2)		
a.	Name any two sources of error in analytical procedures.	2	1	2
b.	Write any two properties of primary standard.	2	1	2
c.	Write the name of the two theories of Indicators.	2	2	2
d.	What is the major limitation of Non aqueous titration?	2	2	6
e.	Which of the following compounds cannot be analysed in the Mohr method? a. NaI b. KCl c. MgBr d. CaCl ₂	2	3	2
f.	EDTA is a dentate ligand.	2	3	2
g.	Write the oxidation number of Mn in KMnO ₄ .	2	4	2
h.	Which is the direct method Iodimetry or Iodometry?	2	4	2
i.	Write the full form of DME.	2	5	2
j.	Give an example of indicator electrode in Potentiometry	2	5	2
Q.2	Attempt any Two Questions	20(2x10)		
a.	Differentiate between primary and secondary standards in pharmaceutical analysis. Elaborate on their importance in the preparation and standardization of solutions.	10	1	2
b.	Define oxidation and reduction concepts in redox titrations. Explore different types of redox titrations and their principles and applications.	10	4	2
c.	Explain the introduction, conductivity cell, conductometric titrations, and their applications.	10	5	2
Q.3	Attempt any Seven Questions	35(7x5)		
a.	Write the steps involved in the gravimetric analysis. Write its advantages and disadvantages.	5	3	2
b.	Write a short note on Metal ion indicators used in complexometry.	5	3	2
c.	Write in detail on limit test for heavy metals.	5	1	2
d.	Write in detail on construction and working of Dropping mercury electrode.	5	5	2
e.	Explain Principal and procedure for the limit test for chloride ions.	5	1	2
f.	Define polarography and explain the Ilkovic equation. Provide an example of a pharmaceutical compound analyzed using polarography.	5	5	2
g.	Write short notes on rotating platinum electrode.	5	5	2
h.	Discuss the significance of primary and secondary standards in pharmaceutical analysis, providing examples of each.	5	1	2
i.	Write a short note on Pharmacopoeia.	5	1	2

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End Semester Examination (December 2024)

B. Pharm. – 1st Semester

Subject Code: BP 103T

Subject Name: Pharmaceutics-I

Time: 3.00 Hour

Total Marks: 75

Q.1	Attempt all Questions	Marks	CO	PO
		20(10*2)		
a.	Some drugs are advised before meals to, 1. Avoid dose dumping. 2. Avoid interaction with food. 3. Improve absorption. 4. Both B & C.		2	2,3
b.	A lesser dose is required if the drug is given by, 1. Oral route. 2. Topical route. 3. Intravenous route. 4. None of the above		2	3,4
c.	Which of the following formulations would not be applicable to ocular administration? 1. Solution 2. Liniment 3. Suspension 4. Ointment		2	2,3
d.	A tablet to treat a headache must first be dissolved in water before swallowing. Which one of the following best describes this type of tablet? 1. Modified release 2. Oral disintegrating 3. Effervescent 4. Buccal		2,3	3,4
e.	Suppository base should _____ at body temperature. 1. Soften 2. Melt 3. Dissolve 4. all of these		4	2,3
f.	The base must be capable of _____ to release its drug for absorption. 1. Softening 2. Melting 3. Dissolving 4. all of these		4	1,2
g.	Which of the following is not an ideal property of bases? 1. It must retain the shape and size 2. It should shrink sufficiently to remove from mould 3. It should be compatible with variety of drugs 4. It should soften or harden on storage		4	2,3
h.	Suppository bases are classified as 1. Fatty or Oily Bases or Oleaginous bases 2. Water Soluble & Water miscible bases Or Hydrophilic bases 3. Emulsifying/Synthetic bases 4. All of these		4	3,1
i.	Name 5 factors affecting posology.		2	1,3
j.	Name 5 headings of label.		1,3	1,2

Q.2	Attempt any Two Questions.	20(2*10)		
a.	Define semisolids, give classification in brief, and discuss mechanisms dermal penetration of drugs		4	3,4
b.	Discuss in detail about stability problems of emulsion and methods to overcome.		3,4	1,2
c.	What are powders? How they classify, Give advantages and disadvantages of Powder dosage forms.		4	2,3
Q.3	Attempt any Seven Questions.	35(7*5)		
a.	Define suspension, its advantages and disadvantages and classifications.		4	2,3
b.	Give detail about Excipients used in semi solid dosage forms		4	3,1
c.	Define gels and give their Evaluation.		4	3,4
d.	Difference between creaming and caking.		4	2,3
e.	Short note on BP.		1	3,4
f.	Write on key features of monophasic dosage forms		2	2,3
g.	Ideal characteristics of solvents used in preparation of liquid dosage forms.		4	3,4
h.	Give advantages and disadvantages of suppositories with its types.		4	2,3
i.	Define Posology and 5 factors which affect selection of dose.		2	1,2

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End Semester Examination (December-2024)
B. Pharm. – 1st Semester

Subject Code: BP101T

Subject Name: Human Anatomy and Physiology I

Time: 3.00 Hour

Total Marks: 75

		Marks	CO	PO
Q.1	Attempt all Questions.	20(10*2)		
a.	Give name of smallest and longest bone of human body.		1	1
b.	Give definition of blood pressure.		2	2
c.	What is the difference between RBC and WBC?		1	1
d.	Give function of bone.		2	2
e.	What is the difference between homeostasis and hemostasis?		1	2
f.	Enlist the name of joints according to movement.		1	1
g.	Which cell of retina known as photoreceptor.		2	2
h.	Give definition of blood.		2	2
i.	Give function of spleen.		2	2
j.	Which nerve are essential for taste and smell.		1	2
Q.2	Attempt any Two Questions.	20(2*10)		
a.	Explain the structure and function of cell.		1	2
b.	Describe the anatomy and function of human skeleton system.		2	2
c.	Define blood pressure. Discuss factors affecting blood pressure and add a note on regulation of blood pressure.		2	2
Q.3	Attempt any Seven Questions.	35(7*5)		
a.	Write a note on anatomy of heart with structure.		2	2
b.	Describe the origin and function of cranial nerves.		2	2
c.	Write a short note on physiology of vision.		1	2
d.	Write a note on spinal nerves.		1	2
e.	Explain the structure and function of lymph node.		1	2
f.	Write a note on blood groups.		1	1
g.	Explain erythropoiesis process.		2	2
h.	Explain the structure and function of skin.		1	2
i.	Discuss physiology of muscle contractions.		2	2

End Semester Examination (December 2024)
B. Pharm. – 1st Semester

Subject Code: BP102T

Subject Name: Pharmaceutical Analysis-I

Time: 3.00 Hour

Total Marks: 75

		Marks	CO	PO
Q.1	Attempt all Questions	20(10x2)		
a.	Name any two sources of error in analytical procedures.	2	1	2
b.	Write any two properties of primary standard.	2	1	2
c.	Write the name of the two theories of Indicators.	2	2	2
d.	What is the major limitation of Non aqueous titration?	2	2	6
e.	Which of the following compounds cannot be analysed in the Mohr method? a. NaI b. KCl c. MgBr d. CaCl ₂	2	3	2
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a.	Write the steps involved in the gravimetric analysis. Write its advantages and disadvantages.	5	3	2
b.	Write a short note on Metal ion indicators used in complexometry.	5	3	2
c.	Write in detail on limit test for heavy metals.	5	1	2
d.	Write in detail on construction and working of Dropping mercury electrode.	5	5	2
e.	Explain Principal and procedure for the limit test for chloride ions.	5	1	2
f.	Define polarography and explain the Ilkovic equation. Provide an example of a pharmaceutical compound analyzed using polarography.	5	5	2
g.	Write short notes on rotating platinum electrode.	5	5	2
h.	Discuss the significance of primary and secondary standards in pharmaceutical analysis, providing examples of each.	5	1	2
i.	Write a short note on Pharmacopoeia.	5	1	2